

**CHELATION OF CHARGED AND UNCHARGED MOLECULES  
WITH PORPHYRIN-BASED COMPOUNDS**

**Maria da Graça Henriques Vicente**

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**[0001]** The benefit of the filing date of provisional application serial number 60/426,612, filed November 15, 2002, is claimed under 35 U.S.C. § 119(e).

**[0002]** This invention pertains to the chelation of charged and uncharged molecules.

**[0003]** Many of the current, commercially available receptors for molecular recognition are difficult and expensive to synthesize, and are typically only obtained in low yield. There is an unfilled need for binding agents that are easy to synthesize, and that are selective for specific substrates, both charged and uncharged.

**[0004]** J. Sessler *et al.*, "Anion binding: a new direction in porphyrin-related research," *Pure & Appl. Chem.*, vol. 65, pp. 393-398 (1993) discloses the use of certain porphyrin or porphyrin-type compounds to chelate certain anionic substrates.

**[0005]** P. Gale *et al.*, "Calix[4]pyrroles: old yet new anion-binding agents," *J. Am. Chem. Soc.*, vol. 118, pp. 5140-5141 (1996) discloses the use of certain octaalkylporphyrinogens as binding agents for certain anionic species.

**[0006]** W. Allen *et al.*, "Binding of neutral substrates by calix[4]pyrroles," *J. Am. Chem. Soc.*, vol. 118, pp. 12471-12472 (1996) discloses the use of certain porphyrin or porphyrin-type compounds to bind certain neutral substrates, both in the solid state and in solution.

**[0007]** J. Sessler *et al.*, "Anion binding: self-assembly of polypyrrolic macrocycles," *Angew. Chem. Int. Ed. Engl.*, vol. 35, pp. 2782-2785 (1996) discloses that certain zwitterionic sapphyrins and certain calix[4]pyrroles can readily self-assemble.

**[0008]** M. Vicente, "Porphyrin-based sensitizers in the detection and treatment of cancer: recent progress," *Curr. Med. Chem.*, vol. 1, pp. 175-194 (2001) provides a review of the use of porphyrins for cancer detection and treatment by photodynamic therapy, boron neutron capture therapy, radiation therapy, and magnetic resonance imaging.

**[0009]** M. Vicente *et al.*, international patent application WO 01/85736 (2001), discloses the use of the same class of porphyrin-based compounds that are used in the present invention, but for a different purpose, namely, in boron neutron capture therapy for treatment of cancer. See also M. Vicente *et al.*, "Syntheses and preliminary biological studies of four *meso*-tetra[(*nido*-carboranylmethyl)phenyl]porphyrins," *Bioorganic & Medicinal Chem.*, vol. 10, pp. 481-492 (2002); A. Maderna *et al.*, "Synthesis of a porphyrin-labelled carboranyl phosphate diester: a potential new drug for boron neutron capture therapy of cancer," *Chem. Commun.*, pp. 1784-1785 (2002); M. Vicente *et al.*, "Syntheses of carbon-carbon linked carboranylated porphyrins for boron neutron capture therapy of cancer," *Tetr. Lett.*, vol. 41, pp. 7623-7627 (2000); and M. Vicente *et al.*, "Synthesis, dark toxicity and induction of in vitro DNA photodamage by a tetra(4-*nido*-carboranyl)porphyrin," *J. Photochemistry and Photobiology*. Vol. 68, pp. 123-132 (2002).

**[0010]** M. Vicente *et al.*, "First Structural Characterization of a Covalently Bonded Porphyrin-Carborane System," *Chem. Commun.*, pp. 483-484 (2001) reported that the large carborane substituents significantly increase the barriers for aryl rotation in tetra(carboranylphenyl)porphyrins.

**[0011]** R. Lauceri *et al.*, "Interactions of Anionic Carboranylated Porphyrins with DNA," *J. Am. Chem. Soc.*, vol. 123, pp. 5835-5836 (2001) reported that the large carborane substituents in tetra(carboranylphenyl)porphyrins influence the basicity of the inner porphyrin nitrogens and their aggregation state.

**[0012]** Additional Citations: R. Barth *et al.*, *Neurosurg.*, vol. 44, pp. 433-451 (1999); T. Dougherty *et al.*, "Photodynamic Therapy," *J. Natl. Cancer Inst.*, vol. 90, pp. 889-905 (1998); and R. Hsi *et al.*, "Photodynamic Therapy in the Treatment of Cancer. Current State of the Art," *Drugs*, vol. 57, pp. 725-734 (1999).

**[0013]** To the inventor's knowledge, no prior reference has suggested that porphyrins containing *nido*-carborane or *closo*-carborane cages would be useful as chelators or specific binding agents.

**[0014]** I have discovered that certain carboranyl-containing porphyrin compounds, such as those disclosed in M. Vicente *et al.*, international patent application WO 01/85736 (2001), bearing chemically stable carbon-carbon bonds between the carborane clusters and the porphyrin ring, may be used to selectively chelate or bind a variety of substrates. The large carborane substituents can significantly increase the steric barriers to aryl rotation in tetra(carboranylphenyl)porphyrins, in addition to influencing the basicity of the inner porphyrin nitrogens and their aggregation state. These carboranylporphyrins are able to effectively bind a variety of molecules and ions by: **(1)** coordination to a pentacoordinated or hexacoordinated metal ion such as Zn(II), Fe(III), Mn(III), Al(III), or Sn(IV) in the macrocyclic core, for example compounds **5, 6, 11, 12, 17, 18, 23, 24, 30, 31, 35, and 36**; **(2)** electrostatic interactions with the carborane groups, for example compounds **3-6, 9-12, 15-18, 21-24, 28-31, and 33-36**; and **(3)** interactions (such as  $\pi$ - $\pi$  interactions) with the aromatic macrocycle, for example compounds **3-6, 9-12, 15-18, 21-24, 28-31, and 33-36**.

**[0015]** Several examples of these interactions have been experimentally observed to date. For example, observed interactions of type (1) include the following: In the solid state, depending on which other components are present during recrystallization (e.g., pyridine, methanol, picric acid, or HCl), zinc(II) carboranylporphyrins have been obtained with different axial ligands: a pyridine, a methanol, a water molecule, or chloride. Certain

molecules have high affinity and selectivity for binding the zinc ion; the carborane groups modulate this specificity. In aqueous solution, both metal-free and Zn(II) complexes of *nido*-carboranylporphyrins have been observed to have an affinity for binding to biomolecules, such as DNA, lipoproteins, and albumin. This binding affinity can strongly influence their biological properties. More generally, binding of type (1) may include, for example, binding to ligands containing electronegative atoms such as nitrogen, oxygen, or sulfur, for example, pyridines, amines, amides, carboxylic acids, hydroxyls, thiols, sulfates, sulfides, and sulfoxides.

**[0016]** UV-Vis spectrophotometry has also shown that *nido*-carboranylporphyrins bind to serum proteins in solution.

**[0017]** Examples of observed interactions of type (2) include the following: The negatively charged *nido*-carboranyl groups have been observed to efficiently bind organic and inorganic cations, for example, sodium, potassium, and piperidinium. These complexes were prepared by ion-exchange using a Dowex 50WX2-100 resin.

**[0018]** Examples of observed interactions of type (3) include the following: Metal-free tetra(*nido*-carboranyl)porphyrins bearing the carborane cages on the *para*-phenyl positions have been observed to protonate at pH ~ 7, while the regioisomers with the *nido*-carborane cages on the *meta*-phenyl positions are essentially unprotonated at pH ~ 7. The protonated carboranylporphyrins are capable of strongly and selectively binding organic and inorganic anions, for example, phosphates, sulfates, sulfonates, carboxylates, thiolates, and alkoxides. A large number of carborane cages (e.g. 4-8, preferably 8) at the periphery of porphyrin macrocycles also make it easy to protonate the porphyrin macrocycle, due to the bulkiness and electron-withdrawing effect of the carborane cages. In the solid state diprotonated octa(*c/oso*-carboranyl)porphyrins have been identified, such as compound **28** (M = 4H), that selectively bind two picrate molecules or two chloride ions.

**[0019]** A surprising interaction of types (1) and (2) has been observed in the solid state for Zn(II) tetra(*c/oso*-carboranyl)porphyrin compound **5**. Close intermolecular contacts were inferred between the BH hydrogens of a carborane group of one porphyrin

and the centrally-chelated zinc ion of another porphyrin. This interaction produced an unusual pseudo-hexacoordinated zinc porphyrin complex. (Zinc porphyrins are usually penta-coordinated).

**[0020]** Negatively-charged *nido*-carboranyporphyrins can efficiently and selectively bind positively charged molecules, as well as neutral and negatively charged molecules, by interaction with the positive macrocyclic core (upon protonation), or by axial ligation to the centrally-chelated metal ions, or both. On the other hand, *closo*-carboranyporphyrins can specifically bind neutral and negatively charged molecules by interaction with the positive macrocyclic core (upon protonation), or by axial ligation to the centrally-chelated metal ions, or both.

**[0021]** The position, nature, and distribution of the bulky carboranyl cages at the periphery of the porphyrin macrocycle affect their selectivity for substrate recognition and binding, both in organic and aqueous solutions, as well as in the solid state. *Closo*- and *nido*-carboranyporphyrins are readily protonated at the inner porphyrin core, and the resulting positively-charged or zwitterionic species can potentially bind anionic, cationic, and polar neutral substrates via electrostatic interactions. Penta- and hexa-coordinated metallo-carboranyporphyrins effectively bind anionic and neutral ligands via axial coordination to the central metal cation. Hydrogen bonding, hydrophobic, and  $\pi$ - $\pi$  interactions can further enhance binding and specificity.

**[0022]** Carboranyporphyrins that may be used in this invention can have varying distributions of carborane cages at the periphery of the macrocycle. The number and distribution of the carborane cages will affect the selectivity and specificity of binding.

**[0023]** When these carboranyporphyrins are complexed to other compounds, the UV-visible and fluorescence spectra of the latter are typically altered, particularly the visible and fluorescence spectra. These carboranyporphyrins will frequently bind more strongly than will the commercially-available, hydrogen-bonding ligands that are typically used today.

**[0024]** Uses for this invention include such things as removing impurities from water, environmental remediation, electronics applications, transport of ionic species generally, sensors for selective ion binding, sensors for selective substrate recognition, separation membranes, electrodes, field effect transistors, and selective binding of specific substrates in biological systems.

**[0025]** Specificity of binding may be modulated by selecting the number of carborane substituents, by selecting *closo*- or *nido*-carborane substituents, by selecting the metal ion in the macrocyclic core, and by selecting the position of substitution for the carborane substituents on the phenyl groups. The carboranes are electron-withdrawing groups, so in general, the more such groups that are present, the more positive charge will tend to be carried by the central metal ion -- although this general tendency can be affected by the distribution and steric hindrance of the carborane groups in a particular compound. A *closo*-carborane cage is uncharged, but electron-withdrawing, while a *nido*-carborane group is negatively charged, and will bind electrostatically to cations on the periphery of the compound. Varying the metal ion in the core can alter its electropositivity, and can alter the affinity for various anions and neutral polar molecules. If the carborane groups are substituted in the *para* position on the phenyl group, then they act primarily as centers of negative charge, but do not tend to form a central cavity. If the carborane groups are substituted in the *meta* position, then they tend to form a cavity that can hold a ligand. In the *meta* position, the phenyl groups are still free to rotate, although the energy barrier to rotation is larger than for the *para*-phenyl substituted carboranylporphyrins. If the carborane groups are substituted in the *ortho* position, then they tend to form a still tighter cavity for the ligand. With *ortho* substitution, the phenyl groups are sterically hindered from rotating, and atropisomers can thereby form (e.g. up, up, up, up -- or up, down, up, down -- or up, up, up, down -- or up, up, down, down). A mixture of atropisomers will generally be the product of syntheses of these compounds; different atropisomers may be separated from one another by means known in the art, e.g., chromatographically.

**[0026]** A *nido*-carborane group is depicted as structure **51** in Fig. 13(a). A *closo*-carborane group is depicted as structure **61** in Fig. 13(b).

**[0027]** Examples of active compounds include the compounds depicted as structures **52**, **53**, and **54** in Figs. 13(c)-(e). In these structures, M may be a metal ion, preferably a pentacoordinated or hexacoordinated metal ion such as Zn(II), Fe(III), Mn(III), Al(III), or Sn(IV), or it may be 2H, the last resulting in a metal-free macrocycle.

**[0028]** The negative charges lie primarily in the boron clusters. The carbon-carbon bonds linking the boron-containing groups to the porphyrin ring make the compounds highly resistant to hydrolysis. By contrast, most prior work with boron-containing porphyrin compounds has relied on ester linkages, which are susceptible to hydrolysis *in vivo*.

**[0029]** These compounds have low toxicity to normal mammalian cells.

**[0030] BRIEF DESCRIPTION OF THE DRAWINGS**

**[0031]** Figure 1 (Scheme 1) depicts the synthesis of compounds **1 – 6**.

**[0032]** Figure 2 (Scheme 2) depicts the synthesis of compounds **7 - 12**.

**[0033]** Figure 3 (Scheme 3) depicts the synthesis of compounds **13 – 18**.

**[0034]** Figure 4 (Scheme 4) depicts the synthesis of compounds **19 - 24**.

**[0035]** Figure 5 (Scheme 5) depicts the synthesis of compounds **25 - 31**.

**[0036]** Figure 6 (Scheme 6) depicts the synthesis of compounds **32- 36**.

**[0037]** Figure 7 depicts the structure inferred from X-ray data for compound **5** binding a methanol molecule axially.

**[0038]** Figure 8 depicts the structure inferred from X-ray data for compound **17** binding a water molecule axially.

**[0039]** Figure 9 depicts the structure inferred from X-ray data for compound **28** binding two chloride ions axially.

**[0040]** Figure 10 depicts the structure inferred from X-ray data for compound **28** binding oxygen atoms from two picrate ions axially.

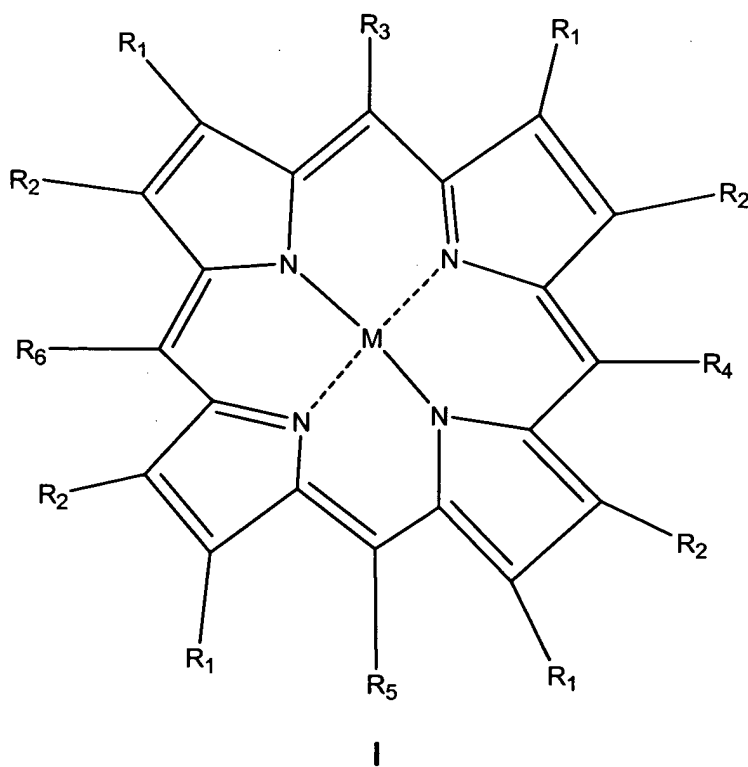
**[0041]** Figure 11 depicts the structure inferred from X-ray data for compound **30** binding a pyridine molecule axially.

**[0042]** Figure 12 depicts the structure inferred from X-ray data for compound **35** binding a pyridine molecule axially.

**[0043]** Figure 13(a) depicts a *nido*-carborane group; and Figure 13(b) depicts a *closo*-carborane group.

**[0044]** Figures 13(c)-(e) depict examples of active compounds.

**[0045]** This invention pertains to the use of certain porphyrin compounds containing carboranyl groups attached to the porphyrin group by a carbon-carbon linkage as chelating agents. As examples, these porphyrin compounds may correspond generally to formula I:

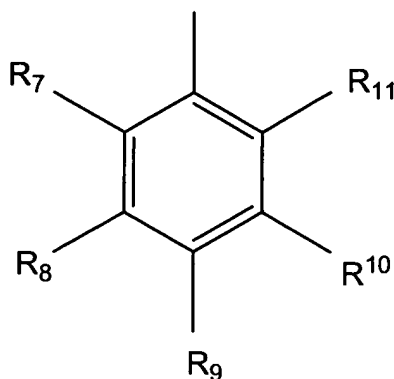




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where M is 2H or a metal ion, preferably a pentacoordinated or hexacoordinated metal ion; R1 and R2 are each independently hydrogen, alkyl or hydroxyalkyl; and R3 through R6 are each independently hydrogen, phenyl, or a substituted phenyl group.

**[0046]** Where one or more of R3 through R6 comprise a substituted phenyl group, that substituted phenyl group corresponds to general formula II:



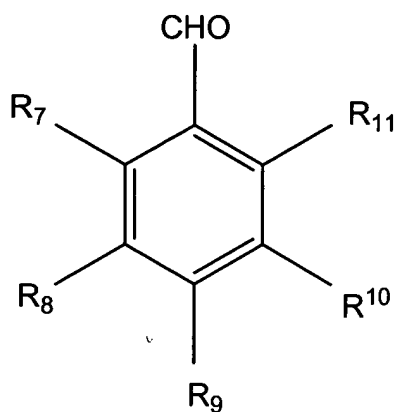
where R7 through R11 are each independently hydrogen or a carboranyl group. The carboranyl group should be attached to the phenyl group by a carbon-carbon linkage (and not, for example, an ether linkage or an ester linkage). Typically, one or two of R7 through R11 are carboranyl groups.

**[0047]** At least one of R3 through R6 should be a substituted phenyl group of general formula II, having at least one carboranyl group attached by a carbon-carbon linkage. More preferably, at least two of R3 through R6 should be a substituted phenyl group of general formula II, each having at least one carboranyl group attached by a carbon-carbon linkage. Most preferably, all of R3 through R6 should be a substituted phenyl group of general formula II, each having at least one carboranyl group attached by a carbon-carbon linkage.

**[0048]** In a preferred embodiment, the compound has from 4 to 8 carboranyl substituents.

**[0049]** The carboranyl-containing porphyrin compounds used in this invention may be synthesized, for example, by the following general method. (Examples of specific

syntheses are described in detail below.) A pyrrole or dipyrrole is reacted with a benzaldehyde using an acid catalyst, such as trifluoroacetic acid. The pyrrole or dipyrrole may be unsubstituted or substituted, for example with alkyl groups. The benzaldehyde has general formula III:



III

where R7 through R11 are each independently hydrogen or a carboranyl group.. The carboranyl group should be attached to the phenyl group by a carbon-carbon linkage (and not, for example, an ether linkage or an ester linkage). Typically, one or two of R7 through R11 are carboranyl groups. The next step is oxidation of the reaction mixture, for example, with tetrachloroquinone or dichlorodicyanobenzoquinone.

**[0050]** Optionally, the synthesis may include complexing the porphyrin compound with a penta- or hexa-coordinated metal ion, e.g., by treating the free base of the porphyrin compound with zinc chloride to form a Zn(II) complex.

**[0051]** Figures 1 through 6 depict, in a few steps, the total synthesis of various carboranylated phenylporphyrins. These *meso*-phenylporphyrin compounds contain carbon-carbon linkages between the carboranyl groups and the porphyrin ring for increased chemical stability. These compounds are soluble in aqueous solution.

**[0052]** In summary, the compounds preferred for use in the present invention contain carbon-carbon linkages between a porphyrin ring and 4-8 carboranyl groups, and are soluble in an aqueous environment.

**[0053] EXAMPLES**

**[0054] Synthesis of Compound 1:**

**[4-(1-methyl-o-carboranyl)methyl]bromobenzene (1):** A two-necked round bottom flask containing 1-methyl-o-carborane (5.00 g, 31.65 mmol) in dry DME (150 mL) was cooled to 0° C under argon. *n*-BuLi (20.0 mL, 1.6 M in hexane) was added dropwise, and the resulting mixture was stirred at 0°C for 30 minutes. A solution of 4-(bromomethyl)bromobenzene (7.91 g, 31.65 mmol) in dry DME (15 mL) was added dropwise. After stirring at 0° C for 10 minutes, the final reaction mixture was warmed to room temperature and then refluxed for 12 hours under argon. The solvent was then removed under vacuum, and the crude solid obtained was purified by recrystallization from dichloromethane/methanol to give the title compound (7.80 g, 75.4% yield) as white crystals. MS (EI) *m/e* 327.1 (*M*<sup>+</sup>); <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.3-3.0 (br, 10H, BH), 2.15 (s, 3H, CH<sub>3</sub>), 3.41 (s, 2H, CH<sub>2</sub>), 7.06 (d, 2H, ArH, *J* = 8.1 Hz), 7.48 (d, 2H, ArH, *J* = 8.1 Hz).

**[0055] Synthesis of Compound 2:**

**[4-(1-methyl-o-carboranyl)methyl]benzaldehyde (2):** A solution of compound 1 (4.00 g, 12.23 mmol) in THF (150 mL) under argon was cooled to -78° C (acetone/dry ice bath). *n*-BuLi (7.6 mL, 1.6 M in hexane) was added dropwise while maintaining the temperature at -78° C. The reaction mixture was stirred for 30 minutes at -78° C, after which dry DMF (5.0 mL, 64.6 mmol) was slowly added. The final mixture was stirred at -78° C for 15 minutes, and then warmed slowly to room temperature. A 2 N HCl solution (150 mL) was added, and the reaction mixture stirred for 2 h at room temperature. The solution was then reduced to a volume of 200 mL and extracted with dichloromethane (4 x 50 mL). The organic extracts were washed once with a saturated aqueous NaHCO<sub>3</sub> solution, and once

with water, and were then dried over anhydrous  $\text{Na}_2\text{SO}_4$ . After removal of the solvent under vacuum, the oily residue was purified by column chromatography on silica gel (dichloromethane/petroleum ether 1:1), yielding the title compound (2.1 g, 62% yield) as a white solid. MS (EI)  $m/e$  276.2 ( $\text{M}^+$ );  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 1.5-3.0 (br, 10H, BH), 2.19 (s, 3H,  $\text{CH}_3$ ), 3.54 (s, 2H,  $\text{CH}_2$ ), 7.38 (d, 2H, ArH,  $J = 8.0$  Hz), 7.89 (d, 2H, ArH,  $J = 8.0$  Hz), 10.04 (s, 1H, CHO).

**[0056]      Synthesis of Compound 3:**

***meso*-tetra[4-(1-methyl-*o*-carboranyl)methylphenyl]porphyrin (3):** A solution of aldehyde **2** (1.16 g, 4.19 mmol) and freshly distilled pyrrole (0.30 mL, 4.32 mmol) in dry dichloromethane (420 mL) was purged with argon for 15 minutes. TFA (0.25 mL, 3.15 mmol) was added to the solution, and the final mixture was stirred at room temperature under argon for 20 hours (at which time the starting compound **2** had essentially disappeared completely, as assayed by TLC). After addition of *p*-chloranil (0.788 g, 3.14 mmol), the final reaction mixture was stirred at room temperature for 2 hours. The solution was concentrated under vacuum to 200 mL, then washed once with a saturated aqueous  $\text{NaHCO}_3$  solution, and once with water, and were then dried over anhydrous  $\text{Na}_2\text{SO}_4$ . The residue after removal of the solvent under vacuum was purified by column chromatography (dichloromethane/petroleum ether 1:1), and the fastest-running porphyrin fraction was collected and recrystallized from dichloromethane/methanol, yielding 0.289 g (21% yield) of the title compound as purple crystals, m.p.  $> 300^\circ\text{C}$ ; MS (MALDI)  $m/e$  1296.0 ( $\text{M}^+$ );  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: -2.80 (br, 2H, NH), 1.6-3.1 (br, 40H, BH), 2.34 (s, 12H,  $\text{CH}_3$ ), 3.81 (s, 8H,  $\text{CH}_2$ ), 7.59 (d, 8H, ArH,  $J = 8.0$  Hz), 8.20 (d, 8H, ArH,  $J = 8.0$  Hz), 8.85 (s, 8H,  $\beta$ -H). UV-Vis ( $\text{CHCl}_3$ )  $\lambda_{\text{max}}$ : 418 nm ( $\epsilon$  467,700), 514 (16,867), 550 (8,132), 590 (5,470), 646 (4,028).

**[0057]      Synthesis of Compound 4:**

**meso-tetra[4-(1-methyl-*nido*-carboranyl)methylphenyl]porphyrin tetrapotassium salt (4):** Porphyrin **3** (0.050 g, 0.0386 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and then stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, the residue was re-dissolved in a 60% acetone aqueous solution, which was then passed slowly over a Dowex 50W2-100 resin in the potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution, and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.051 g (94 % yield) of the title compound, m.p. > 300° C. <sup>1</sup>H-NMR (CD<sub>3</sub>COCD<sub>3</sub>) δ ppm: -2.70 (s, 2H, NH), -2.45 to -1.90 (br, 4H, BH), 0.9-2.4 (br, 32H, BH), 1.59 (s, 12H, CH<sub>3</sub>), 3.50 (s, 8H, CH<sub>2</sub>), 7.81 (d, 8H, ArH, J = 8.0 Hz), 8.08 (d, 8H, ArH, J = 8.0 Hz), 8.90 (s, 8H, β-H). UV-Vis (acetone) λ<sub>max</sub>: 420 nm (ε 349,700), 516 (13,595), 554 (12,410), 594 (4,130), 650 (5,990).

**[0058]      Synthesis of Compound 5:**

**Zn(II)-meso-tetra[4-(1-methyl-*o*-carborane)methylphenyl]porphyrin (5):** To a solution of porphyrin **3** (0.150 g) in dichloromethane (150 mL), THF (10 mL), and pyridine (0.5 mL) were added with ZnCl<sub>2</sub>•2H<sub>2</sub>O, and the final mixture was stirred at room temperature under argon overnight. The mixture was then washed once with water, dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>, and the solvent evaporated under vacuum. The residue was purified by column chromatography (dichloromethane/petroleum ether 1:1.5), and the pink-colored fraction was collected and recrystallized from dichloromethane/methanol to give 0.135 g (92 % yield) of the title compound as purple crystals, m.p. > 300° C; MS m/e 1358.6; <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.6-3.0 (br, 40H, BH), 2.33 (s, 12H, CH<sub>3</sub>), 3.80 (s, 8H, CH<sub>2</sub>), 7.57 (br s, 8H, ArH), 8.19 (br s, 8H, ArH), 8.95 (br s, 8H, β-H). UV-Vis (CHCl<sub>3</sub>) λ<sub>max</sub>: 424 nm (ε 577,000), 554 (20,102), 596 (6,380).

**[0059]      Synthesis of Compound 6:**

**Zn(II)-*meso*-tetra[4-(1-methyl-*nido*-carboranyl)methylphenyl]porphyrin tetrapotassium salt (6):** The Zn(II) complex **5** (0.075 g, 0.055 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, and the residue was re-dissolved in a 60% aqueous acetone solution and passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution, and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.078 g (96 % yield) of the title compound, m.p. > 300° C. <sup>1</sup>H-NMR (CD<sub>3</sub>COCD<sub>3</sub>) δ ppm: -2.48 to -1.95 (br, 4H, BH), 0.9-2.4 (br, 32H, BH), 1.59 (s, 12H, CH<sub>3</sub>), 3.50 (s, 8H, CH<sub>2</sub>), 7.77 (d, 8H, ArH, J = 8.0 Hz), 8.11 (d, 8H, ArH, J = 8.0 Hz), 8.92 (s, 8H, β-H). UV-Vis (acetone) λ<sub>max</sub>: 422 nm (ε 479,000), 554 (13,870), 596 (6,595).

**[0060]      Synthesis of Compound 7:**

**[3-(1-methyl-*o*-carboranyl)methyl]bromobenzene (7):** A two-necked round bottom flask containing 1-methyl-*o*-carborane (3.00 g, 18.99 mmol) in dry THF (150 mL) was cooled to 0° C under argon. *n*-BuLi (12.0 mL, 1.6 M in hexane) was added dropwise, and the resulting mixture was stirred at 0° C and then cooled to -10° C. A solution of anhydrous LiI (0.350 g, 2.61 mmol) in THF (2.5 mL) was added, followed by a solution of 3-(bromomethyl)bromobenzene (5.00 g, 20.00 mmol) in THF (10 mL). After stirring at -10° C for 15 minutes, the final reaction mixture was warmed to room temperature and stirred for 12 hours under argon. The reaction mixture was then washed with water (2 x 25 mL), extracted with diethyl ether (3 x 25 mL), and dried over Na<sub>2</sub>SO<sub>4</sub>. The solvent was then removed under vacuum, and the resulting crude solid was purified by column chromatography (silica gel, dichloromethane/petroleum ether 1:9) to give the title compound (4.25 g, 65.0% yield). <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.3-3.1 (br, 10H, BH), 2.16 (s,

3H, CH<sub>3</sub>), 3.42 (s, 2H, CH<sub>2</sub>), 7.13 (d, 1H, ArH, J = 7.8 Hz), 7.23 (t, 1H, ArH, J = 7.8 Hz), 7.33 (s, 1H, ArH), 7.47 (d, 1H, ArH, J = 7.8 Hz).

**[0061]      Synthesis of Compound 8:**

**[3-(1-methyl-o-carboranyl)methyl]benzaldehyde (8):** A solution of compound 7 (1.00 g, 3.06 mmol) in THF (25 mL) under argon was cooled to –78° C (acetone/dry ice bath). *n*-BuLi (2.0 mL, 1.6 M in hexane) was added dropwise while maintaining the temperature at –78° C. The reaction mixture was stirred for 30 minutes at –78° C before dry DMF (1.0 mL, 17.5 mmol) was slowly added. The final mixture was stirred at –78° C for 15 minutes, and was then warmed slowly to room temperature. A 2 N HCl solution (25 mL) was added, and the reaction mixture was stirred for 2 h at room temperature. The solution was then reduced to a volume of 200 mL and extracted with dichloromethane (4 x 50 mL). The organic extracts were washed once with a saturated aqueous NaHCO<sub>3</sub> solution, and once with water, and were then dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After removal of the solvent under vacuum, the oily residue was purified by column chromatography on silica gel (dichloromethane/petroleum ether 1:1), yielding the title compound (0.668 g, 79.1% yield) as a white solid. <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.4-3.1 (br, 10H, BH), 2.19 (s, 3H, CH<sub>3</sub>), 3.55 (s, 2H, CH<sub>2</sub>), 7.48 (d, 1H, ArH, J = 7.8 Hz), 7.56 (t, 1H, ArH, J = 7.8 Hz), 7.70 (s, 1H, ArH), 7.85 (d, 1H, ArH, J = 7.8 Hz), 10.04 (s, 1H, CHO).

**[0062]      Synthesis of Compound 9:**

**meso-tetra[3-(1-methyl-o-carboranyl)methylphenyl]porphyrin (9):** A solution of aldehyde 8 (0.660 g, 2.39 mmol) and freshly distilled pyrrole (0.18 mL, 2.59 mmol) in dry dichloromethane (240 mL) was purged with argon for 45 minutes. TFA (0.15 mL, 1.89 mmol) was added to the solution, and the final mixture was stirred at room temperature under argon for 18 hours. After addition of *p*-chloranil (0.440 g, 1.77 mmol), the final reaction mixture was stirred at room temperature for 3 hours. The organic solution was washed once with a saturated aqueous NaHCO<sub>3</sub> solution, and once with water, and were



then dried over anhydrous  $\text{Na}_2\text{SO}_4$ . The residue obtained after removal of the solvent under vacuum was purified by column chromatography (dichloromethane/petroleum ether 1:1), and the porphyrin fraction was collected and recrystallized from dichloromethane/methanol, yielding 0.252 g (33% yield) of the title compound as purple crystals, m.p.  $> 300^\circ\text{C}$ ; MS (MALDI)  $m/e$  1296.0 ( $\text{M}^+$ );  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: -2.84 (br, 2H, NH), 1.5-3.0 (br, 40H, BH), 2.20 (s, 12H,  $\text{CH}_3$ ), 3.74 (s, 8H,  $\text{CH}_2$ ), 7.62 (d, 4H, ArH), 7.74 (d, 4H, ArH), 8.05 (d, 4H, ArH), 8.18 (d, 4H, ArH), 8.84 (s, 8H,  $\beta\text{-H}$ ). UV-Vis ( $\text{CHCl}_3$ )  $\lambda_{\text{max}}$ : 419 nm, 516, 548, 590, 646.

**[0063]      Synthesis of Compound 10:**

***meso-tetra[3-(1-methyl-*nido*-carboranyl)methylphenyl]porphyrin tetrapotassium salt (10)***: Porphyrin **9** (0.049 g, 0.0378 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, the residue re-dissolved in a 60% aqueous acetone solution, and was then passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution, and was again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.050 g (94 % yield) of the title compound, m.p.  $> 300^\circ\text{C}$ . UV-Vis (acetone)  $\lambda_{\text{max}}$ : 431 nm, 511, 546, 590, 647.

**[0064]      Synthesis of Compound 13:**

**4-Ethynylbenzaldehyde (13)**: To a solution of 4-bromobenzaldehyde (10.00 g, 54.08 mmol) and triphenylphosphine (0.500 g, 1.91 mmol) in anhydrous triethylamine (80 mL) under argon were added ethynyltrimethylsilane (6.00 g, 61.09 mmol) and palladium (II) acetate (0.100g, 0.445 mmol). The final mixture was heated to reflux for 2 hours, and was then cooled to room temperature and filtered. The filtrate was concentrated under vacuum to a thick oil, which was purified by column chromatography (dichloromethane/petroleum ether 1:4) and recrystallized from cold cyclohexane to give 10.5 g (96.1% yield) of 4-

(trimethylsilylethynyl)benzaldehyde [MS  $m/e$  187.2 ( $M^+$ );  $^1H$ -NMR ( $CDCl_3$ )  $\delta$  ppm: 0.27 (s, 9H,  $SiMe_3$ ), 7.60 (d, 2H, ArH,  $J = 8.1$  Hz), 7.82 (d, 2H, ArH,  $J = 8.1$  Hz), 10.00 (s, 1H, CHO)]. This compound (8.00 g, 39.59 mmol) was treated with  $K_2CO_3$  (0.500 g) in methanol (50 mL) at 25° C for 2 hours under Argon. The solvent was removed under vacuum, and the residue was dissolved in dichloromethane (100 mL). This solution was washed once with a saturated aqueous  $NaHCO_3$  solution, and once with water, was then dried over anhydrous  $Na_2SO_4$ , and the solvent was evaporated under vacuum. The yellow residue was purified by column chromatography using dichloromethane/petroleum ether 1:4, and then recrystallized from cold cyclohexane to give 4.40 g (85.5% yield) of the title compound; MS (EI)  $m/e$  130.0; ( $M^+$ ).  $^1H$ -NMR ( $CDCl_3$ )  $\delta$  ppm: 3.30 (s, 1H, CH), 7.64 (d, 2H, ArH,  $J = 8.1$  Hz), 7.84 (d, 2H, ArH,  $J = 8.1$  Hz), 10.02 (s, 1H, CHO).

**[0065]      Synthesis of Compound 14:**

**4-(*o*-carboranyl)benzaldehyde (14):**  $BF_3 \cdot OEt_2$  (0.654 g, 4.62 mmol) was added under argon to a 0° C solution of 4-ethynylbenzaldehyde (**13**) (6.00 g, 46.15 mmol) and 1,2-ethanedithiol (5.00 g, 53.09 mmol). The mixture was stirred at room temperature under argon for 15 minutes. The reaction mixture was then washed once with a 10% aqueous NaOH solution, and once with a saturated aqueous NaCl solution, before being dried over anhydrous  $Na_2SO_4$ . The solvent was then evaporated under vacuum. Purification of the resulting residue by column chromatography (dichloromethane/petroleum ether 1:4) gave *p*-ethynylbenzyl(1,3-dithiane) (7.5 g, 79 % yield) as a yellow solid [MS (EI)  $m/e$  206.0 ( $M^+$ );  $^1H$ -NMR ( $CDCl_3$ )  $\delta$  ppm: 3.07 (s, 1H, CH), 3.38 and 3.51 (m, 2H each,  $CH_2CH_2$ ), 5.61 (s, 1H, SCH), 7.42 (d, 2H, ArH,  $J = 8.1$  Hz), 7.47 (d, 2H, ArH,  $J = 8.1$  Hz)]. Decaborane (3.00 g, 24.59 mmol), ethyl sulfide (5.00 g, 55.44 mmol), and dry toluene (50 mL) were combined in a Schlenk tube equipped with a stir bar. This solution was heated to 40° C for 3 hours and to 60° C for 2 hours, and was then allowed to cool to room temperature. To this mixture was added a solution of *p*-ethynylbenzyl(1,3-dithiane) (5.00 g, 24.26 mmol) in dry toluene (10 mL). The final reaction mixture was slowly warmed to 80° C, and was held at this temperature and stirred for 3 days. After then cooling to room temperature, the mixture

was concentrated under vacuum and the resulting oil was dissolved in methanol (250 mL) and heated to reflux until liberation of hydrogen ceased (approximately 60 minutes). At room temperature a 50% aqueous HCl solution (2 to 3 mL) was cautiously added, and the mixture was again heated to reflux until hydrogen evolution was complete (approximately 30 minutes). After cooling to room temperature, the reaction mixture was diluted with ethanol, and excess ethyl sulfide was removed by ethanol-ethyl sulfide co-distillation. The remaining residue was concentrated under vacuum. To a solution of the resulting residue in benzene (100 mL) at 5° C was added 100 mL of a cold 10% aqueous NaOH solution, and the final mixture was stirred vigorously for 15 minutes. The organic layer was separated, washed with water (3 x 25 mL), and dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After evaporation of the solvent, the residue obtained was purified by column chromatography (dichloromethane/petroleum ether 1:4), yielding 5.25 g (66.8 % yield) of *p*-(*o*-carboranyl)benzyl(1,3-dithiane) [MS (EI) *m/e* 324.1 (M<sup>+</sup>); <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.6-3.3 (br, 10H, BH), 3.36 and 3.47 (m, 2H each, CH<sub>2</sub>CH<sub>2</sub>), 3.91 (br s, 1H, *o*-carborane-CH), 5.56 (s, 1H, SCH), 7.40 (d, 2H, ArH, *J* = 8.1 Hz), 7.46 (d, 2H, ArH, *J* = 8.1 Hz)]. To a solution of this compound (5.00 g, 15.43 mmol) in 5% aqueous THF (25 mL) under argon was added dropwise a solution of HgClO<sub>4</sub> (12.50 g, 31.29 mmol) in THF (15 mL). The final mixture was stirred at room temperature for 15 minutes before being filtered, and the precipitate was washed 3 times with 25 mL diethyl ether. The filtrate was washed with a saturated aqueous Na<sub>2</sub>CO<sub>3</sub> solution (3 x 25 mL) and with water (2 x 25 mL), before being dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. The residue remaining after evaporation of the solvent was purified by column chromatography (dichloromethane/petroleum ether 1:4) to give the title compound (3.27 g, 85.6 % yield); MS (EI) *m/e* 248.2 (M<sup>+</sup>); <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.60 – 3.2 (br, 10H, BH), 4.03 (br s, 1H, *o*-carborane-CH), 7.65 (d, 2H, ArH, *J* = 8.4 Hz), 7.86 (d, 2H, ArH, *J* = 8.4 Hz), 10.04 (s, 1H, CHO).

**[0066]      Synthesis of Compound 15:**

***meso*-tetra[4-(*o*-carboranyl)phenyl]porphyrin (15):** A solution of aldehyde **14** (1.05 g, 4.23 mmol) and freshly distilled pyrrole (0.30 mL, 4.32 mmol) in dry dichloromethane (430

mL) was purged with argon for 30 minutes. TFA (0.20 mL, 2.52 mmol) was added to the solution, and the final mixture was stirred at room temperature under argon for 24 hours. After addition of *p*-chloranil (0.780 g, 3.14 mmol) the final reaction mixture was stirred at room temperature for 3 hours. The solution was concentrated under vacuum to 300 mL, then washed once with a saturated aqueous NaHCO<sub>3</sub> solution, and once with water before being dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After evaporation of the solvent under vacuum, the resulting residue was purified by column chromatography (dichloromethane/petroleum ether 1:2), and the fastest-running porphyrin fraction was collected and recrystallized from dichloromethane/methanol, yielding 0.220 g (17.7% yield) of the title compound as purple crystals, m.p. > 300 ° C; MS (MALDI) m/e 1184.5 (M+1); <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: -2.89 (br, 2H, NH), 1.7-3.5 (br, 40H, BH), 4.28 (br s, 4H, *o*-carborane-CH), 7.89 (d, 8H, ArH, J = 8.0 Hz), 8.17 (d, 8H, ArH, J = 8.0 Hz), 8.78 (s, 8H, β-H). <sup>1</sup>H-NMR (d-TFA/CDCl<sub>3</sub>) δ ppm: -0.97 (br, 4H, NH), 1.8-3.4 (br, 40H, BH), 4.31 (br s, 4H, *o*-carborane-CH), 8.13 (d, 8H, J = 8.0 Hz), 8.51 (d, 8H, J = 8.0 Hz), 8.68 (s, 8H, β-H). UV-Vis (CHCl<sub>3</sub>) λ<sub>max</sub>: 418 nm (ε 464,700), 514 (17,165), 550 (8,300), 590 (5,635), 646 (4,035).

**[0067]      Synthesis of Compound 16:**

***meso*-tetra[4-(*nido*-carboranyl)phenyl]porphyrin tetrapotassium salt (16):** Porphyrin **15** (0.0500 g, 0.0423 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, the residue was re-dissolved in a 60% aqueous acetone solution, and was passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution, and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.0494 g (90.2% yield) of the title compound, m.p. > 300° C. <sup>1</sup>H-NMR (CD<sub>3</sub>COCD<sub>3</sub>) δ ppm: -2.78 (s, 2H, NH), -2.45 to -1.90 (br, 4H, BH), 0.8-2.4 (br, 32H, BH), 2.57 (br s, 4H, *nido*-carborane-CH), 7.66 (d, 8H, ArH, J = 8.0 Hz), 7.97 (d, 8H,

ArH,  $J = 8.0$  Hz), 8.87 (s, 8H,  $\beta$ -H). UV-Vis (acetone)  $\lambda_{\text{max}}$ : 420 nm ( $\epsilon$  302,900), 516 (11,560), 554 (10,580), 594 (3,335), 650 (4,875).

**[0068]      Synthesis of Compound 17:**

**Zn(II)-meso-tetra[4-(*o*-carboranyl)phenyl]porphyrin (17):** To a solution of porphyrin **15** (0.085 g, 0.072 mmol) in dichloromethane (50 mL), THF (4.0 mL), and pyridine (0.5 mL) was added  $\text{ZnCl}_2 \cdot 2\text{H}_2\text{O}$  (0.30 g), and the final mixture was stirred at room temperature under argon overnight. The mixture was then washed once with water, dried over anhydrous  $\text{Na}_2\text{SO}_4$ , and the solvent was evaporated under vacuum. The residue was purified by column chromatography (dichloromethane/cyclohexane 1:2), and the pink-colored fraction was collected and recrystallized from dichloromethane/methanol to give 0.085 g (94.7% yield) of the title compound as purple crystals, m.p.  $> 300^\circ\text{C}$ ; MS (MALDI)  $m/e$  1246.7 ( $\text{M}^+$ );  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 1.6-3.6 (br, 40H, BH), 4.30 (br s, 4H, *o*-carborane-CH), 7.91 (br s, 8H, ArH), 8.17 (br s, 8H, ArH), 8.88 (br s, 8H,  $\beta$ -H). UV-Vis ( $\text{CH}_2\text{Cl}_2$ )  $\lambda_{\text{max}}$ : 424 nm ( $\epsilon$  607,400), 554 (22,566), 594 (6,781).

**[0069]      Synthesis of Compound 18:**

**Zn(II)-meso-tetra[4-(*nido*-carboranyl)phenyl]porphyrin tetrapotassium salt (18):** The Zn(II) complex **17** (0.0600 g, 0.0481 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, and the residue was re-dissolved in a 60% aqueous acetone solution and passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.0615 g (94.0% yield) of the title compound;  $^1\text{H-NMR}$  ( $\text{CD}_3\text{COCD}_3$ )  $\delta$  ppm: -2.54 to -1.78 (br, 4H, BH), 0.6-2.2 (br, 32H, BH), 2.58 (br s, 4H, *nido*-

carborane-CH), 7.64 (d, 8H, ArH,  $J = 8.1$  Hz), 7.95 (d, 8H, ArH,  $J = 8.1$  Hz), 8.88 (s, 8H,  $\beta$ -H). UV-Vis (acetone)  $\lambda_{\text{max}}$ : 426 nm ( $\epsilon$  432,000), 558 (14,380), 598 (9,513).

**[0070]      Synthesis of Compound 19:**

**3-Ethynylbenzaldehyde (19):** To a solution of 3-bromobenzaldehyde (10.00 g, 54.08 mmol) and triphenylphosphine (0.500 g, 1.91 mmol) in anhydrous triethylamine (100 mL) under argon were added ethynyltrimethylsilane (6.00 g, 61.09 mmol) and palladium (II) acetate (0.100g, 0.445 mmol). The final mixture was heated to reflux for 2 hours, and was then cooled to room temperature and filtered. The filtrate was concentrated under vacuum to a thick oil, which was purified by column chromatography (dichloromethane/petroleum ether 1:4) to give 8.52 g (78.0% yield) of 3-(trimethylsilylethynyl)benzaldehyde [ $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 0.26 (s, 9H,  $\text{SiMe}_3$ ), 7.47 (t, 1H, ArH,  $J = 7.5$  Hz), 7.70 (d, 1H, ArH,  $J = 7.5$  Hz), 7.82 (d, 1H, ArH,  $J = 7.5$  Hz), 7.96 (s, 1H ArH), 9.98 (s, 1H, CHO)]. This compound (5.00 g, 24.74 mmol) was treated with  $\text{K}_2\text{CO}_3$  (0.500 g) in methanol (50 mL) at 25° C for 2 hours under argon. The solvent was removed under vacuum, and the residue was dissolved in dichloromethane (100 mL). This solution was washed once with a saturated aqueous  $\text{NaHCO}_3$  solution, and once with water, before being dried over anhydrous  $\text{Na}_2\text{SO}_4$  and the solvent evaporated under vacuum. The yellow residue was purified by column chromatography using dichloromethane/petroleum ether 1:4 for elution, and recrystallized from cyclohexane to give 2.80 g (87.1% yield) of the title compound; MS (EI)  $m/e$  130.0; ( $\text{M}^+$ ).  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 3.20 (s, 1H, CH), 7.51 (t, 1H, ArH,  $J = 7.8$  Hz), 7.74 (d, 1H, ArH,  $J = 7.8$  Hz), 7.87 (d, 1H, ArH,  $J = 7.8$  Hz), 7.99 (s, 1H, ArH), 10.02 (s, 1H, CHO).

**[0071]      Synthesis of Compound 20:**

**3-(o-carboranyl)benzaldehyde (20):**  $\text{BF}_3 \cdot \text{OEt}_2$  (0.11 g, 0.77 mmol) was added at 0° C under argon to a solution of 3-ethynylbenzaldehyde (**19**) (1.00 g, 7.69 mmol) and 1,2-ethanedithiol (0.73 g, 7.75 mmol). This mixture was stirred at room temperature under

argon for 15 minutes. The reaction mixture was then washed once with a 10% aqueous NaOH solution, once with an aqueous saturated NaHCO<sub>3</sub> solution, and once with water before it was dried over anhydrous Na<sub>2</sub>SO<sub>4</sub> and the solvent was evaporated under vacuum. Purification of the resulting residue by column chromatography (dichloromethane/petroleum ether 1:4) gave 1.28 g (80.7% yield) of *m*-ethynylbenzyl(1,3-dithiane) [MS (EI) *m/e* 206.0 (*M*<sup>+</sup>); <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 3.07 (s, 1H, CH), 3.37 and 3.50 (m, 2H each, CH<sub>2</sub>CH<sub>2</sub>), 5.59 (s, 1H, SCH), 7.27 (t, 1H, ArH, *J* = 7.8 Hz), 7.38 (d, 1H, ArH, *J* = 7.8 Hz), 7.50 (d, 1H, ArH, *J* = 7.8 Hz), 7.66 (s, 1H, ArH)]. Decaborane (0.0500 g, 4.10 mmol), ethyl sulfide (0.750 g, 8.32 mmol) and dry toluene (25 mL) were combined in a Schlenk tube equipped with a stir bar. This solution was heated to 40° C for 3 hours and to 60° C for 2 hours, and was then allowed to cool to room temperature. To this mixture was added a solution of *m*-ethynylbenzyl(1,3-dithiane) (0.800 g, 3.88 mmol) in dry toluene (5 mL), and the final reaction mixture was slowly warmed to 80° C, and was stirred at this temperature for 3 days. After cooling to room temperature, the mixture was concentrated under vacuum and the resulting oil was dissolved in methanol (100 mL) and heated to reflux until liberation of hydrogen subsided (approximately 60 minutes). At room temperature a 50% aqueous HCl solution (1.0 mL) was cautiously added, and the mixture was again heated to reflux until hydrogen evolution was complete (approximately 30 minutes). After cooling to room temperature, the reaction mixture was diluted with ethanol, and the excess ethyl sulfide was removed by ethanol-ethyl sulfide co-distillation. The remaining residue was concentrated under vacuum. To a solution of the resulting residue in benzene (50 mL) at 5° C was added 100 mL of a cold 10% aqueous NaOH solution, and the final mixture was stirred vigorously for 15 minutes. The organic layer was separated, washed with water (3 x 15 mL) and dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After evaporation of the solvent, the resulting residue was purified by column chromatography (dichloromethane/petroleum ether 1:4), yielding 0.905 g (72.0 % yield) of *m*-(*o*-carboranyl)benzyl(1,3-dithiane) [<sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.40-3.20 (br, 10H, BH), 3.39 and 3.49 (m, 2H each, CH<sub>2</sub>CH<sub>2</sub>), 3.97 (br s, 1H, *o*-carborane-CH), 5.58 (s, 1H, SCH), 7.28 (t, 1H, ArH, *J* = 7.8 Hz), 7.39 (d, 1H, ArH, *J* = 7.8 Hz), 7.55 (d, 1H, ArH, *J* = 7.8 Hz), 7.63 (s,

1H, ArH)]. To a solution of the latter compound (1.00 g, 3.09 mmol) in 5% aqueous THF (10 mL) under argon was added dropwise a solution of  $\text{HgClO}_4$  (2.50 g, 6.26 mmol) in THF (5.0 mL). The final mixture was stirred at room temperature for 15 minutes before being filtered, and the precipitate was washed with 25 mL of diethyl ether. The filtrate was then washed with a saturated aqueous  $\text{Na}_2\text{CO}_3$  solution (3 x 10 mL) and with water (2 x 10 mL), before being dried over anhydrous  $\text{Na}_2\text{SO}_4$ . The residue remaining after evaporation of the solvent was purified by column chromatography (dichloromethane/petroleum ether 1:4) to give the title compound (0.677 g, 88.5 % yield);  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 1.5 – 3.3 (br, 10H, BH), 4.04 (br s, 1H, o-carborane-CH), 7.56 (t, 1H, ArH,  $J = 7.8$  Hz), 7.79 (d, 1H, ArH,  $J = 7.8$  Hz), 7.91 (d, 1H, ArH,  $J = 7.8$  Hz), 7.96 (s, 1H, ArH), 10.02 (s, 1H, CHO).

**[0072]      Synthesis of Compound 21:**

**meso-tetra[3-(o-carboranyl)phenyl]porphyrin (21):** A solution of aldehyde **20** (0.702 g, 2.83 mmol) and freshly distilled pyrrole (0.200 mL, 2.88 mmol) in dry dichloromethane (285 mL) was purged with argon for 30 minutes. TFA (0.100 mL, 1.26 mmol) was added to the solution, and the final mixture was stirred at room temperature under argon for 18 hours. After addition of *p*-chloranil (0.522 g, 2.10 mmol) the final reaction mixture was stirred at room temperature for 3 hours. The solution was concentrated under vacuum to 200 mL, then washed once with water, once with a saturated aqueous  $\text{NaHCO}_3$  solution, and once again with water before being dried over anhydrous  $\text{Na}_2\text{SO}_4$ . After evaporation of the solvent under vacuum, the resulting residue was purified by column chromatography (dichloromethane/petroleum ether 1:2) and the fastest-running porphyrin fraction was collected and recrystallized from dichloromethane/methanol, yielding 0.140 g (16.7% yield) of the title compound as purple crystals; MS (MALDI)  $m/e$  1184.  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: -2.88 (br, 2H, NH), 1.6-3.5 (br, 40H, BH), 4.19 (br s, 4H, o-carborane-CH), 7.78 (m, 4H, ArH), 7.94 (m, 4H, ArH), 8.27 (m, 4H, ArH), 8.33 (m, 4H, ArH), 8.80 (s, 8H,  $\beta$ -H).



**[0073]      Synthesis of Compound 22:**

**meso-tetra[3-(nido-carboranyl)phenyl]porphyrin tetrapotassium salt (22):**

Porphyrin **21** (0.010 g, 0.008 mmol) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and the mixture was stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, and the residue was re-dissolved in a 60% aqueous acetone solution and then passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.0108 g (98.1% yield) of the title compound, m.p. > 300°C. <sup>1</sup>H-NMR (CD<sub>3</sub>COCD<sub>3</sub>) δ ppm -2.70 (s, 2H, NH), -2.40 to -1.90 (br, 4H, BH), 0.8-2.3 (br, 32H, BH), 2.48 (br s, 4H, *nido*-carborane-CH), 7.49 (m, 4H, ArH), 7.65 (m, 4H, ArH), 7.84 (m, 4H, ArH), 8.15 (m, 4H, ArH), 8.88 (s, 8H, β-H). UV-Vis (acetone) λ<sub>max</sub>: 416 nm (ε 326,300), 512 (14,300), 547 (8,000), 590 (4,300), 646 (4,300).

**[0074]      Synthesis of Compound 25:**

**Bis-(3,5-bromomethyl)bromobenzene (25):** To a refluxing solution of 3,5-dimethylbromobenzene (4.63 g) in dry CCl<sub>4</sub> (300 mL) under argon were added NBS (9.79 g) and benzoyl peroxide (0.80 g) in portions over a one hour period. The final reaction mixture was refluxed with stirring under argon for 16 hours. After cooling to room temperature, the reaction mixture was filtered, and the filtrate was washed once with a saturated aqueous NaHCO<sub>3</sub> solution and once with water. The organic solution was dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>, and the solvent was evaporated under vacuum. The resulting residue was purified by column chromatography using dichloromethane/petroleum ether 1:9, and the main product was collected and recrystallized from n-hexane, yielding 2.83 g (33% yield) of the title compound; <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 4.40 (s, 2H, CH<sub>2</sub>), 7.34 (s, 1H), 7.47 (s, 2H).

**[0075]      Synthesis of Compound 26:**

**Bis[3,5-(1-methyl-o-carboranyl)methyl]bromobenzene (26):** *n*-BuLi (5.2 mL, 1.6 M in hexane) was added dropwise to a solution of methyl-o-carborane (1.39 g, 8.80 mmol) in dry THF (80 mL), at a temperature between -5° and 0° C under argon. The mixture was stirred at this temperature range for 90 minutes, and then cooled to -15° to -20° C (ice/salt bath). A solution of Lil (0.166 g, 1.27 mmol) in dry THF (15 mL) and compound **25** (1.372 g, 4.00 mmol) was added, and the final reaction mixture was allowed to warm to room temperature and was stirred for 16 hours. After the reaction was quenched with water, the resulting mixture was extracted with diethyl ether. The organic extracts were washed once with water and once with brine, were dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>, and the solvent was removed under vacuum. The crude product was purified by column chromatography using dichloromethane/petroleum ether 2:8 for elution, and the main product was collected and recrystallized from *n*-hexane to give 1.26 g (63% yield) of the title compound; MS *m/e* 497.3; <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.4-3.0 (br, 20H, BH), 2.17 (s, 6H, CH<sub>3</sub>), 3.43 (s, 4H, CH<sub>2</sub>), 6.96 (s, 1H), 7.31 (s, 2H).

**[0076]      Synthesis of Compound 27:**

**Bis[3,5-(1-methyl-o-carboranyl)methyl]benzaldehyde (27):** A solution of compound **26** (0.994 g) in THF (20 mL) under argon was cooled to -78° C. *n*-BuLi (1.4 mL, 1.6 M in hexane) was added dropwise via syringe. After the reaction mixture was stirred for 30 minutes at -78° C, dry DMF (0.77 mL) was slowly added. The resulting yellow mixture was stirred at -78° C for 30 minutes, and was then warmed to 0° C and stirred at this temperature for one hour. A 5% aqueous HCl solution was added until the pH of the reaction mixture was between 2 and 3, and the final mixture was then stirred at room temperature. The aqueous layer was extracted 4 times with diethyl ether, the organic fraction was dried over anhydrous MgSO<sub>4</sub>, and the solvent was evaporated under vacuum. Purification by column chromatography (dichloromethane/petroleum ether 2:3), afforded the title compound (0.632 g) in 70.9% yield; MS *m/e* 446.4; <sup>1</sup>H-NMR (CDCl<sub>3</sub>) δ ppm: 1.5-3.0

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(br, 20H, BH), 2.20 (s, 6H, CH<sub>3</sub>), 3.55 (s, 4H, CH<sub>2</sub>), 7.30 (d, 1H, J = 1.6 Hz), 7.67 (d, 2H, J = 1.6 Hz), 10.03 (s, 1H, CHO).

**[0077]      Synthesis of Compound 28:**

**meso-tetra[bis-3,5-(1-methyl-o-carboranyl)methylphenyl]porphyrin (28):** A solution of aldehyde **27** (0.243 g, 0.54 mmol) and freshly distilled pyrrole (0.050 mL g, 0.72 mmol) in dry dichloromethane (55 mL) was purged with argon for 15 minutes. TFA was added, and the final solution was stirred at room temperature overnight (until the starting aldehyde had completely disappeared, and 2 new spots had formed as assayed by TLC). After oxidation with *p*-chloranil (0.102 g, 0.41 mmol) for 6 hours at room temperature, the final reaction mixture was washed once with a saturated aqueous NaHCO<sub>3</sub> solution and once with water before being dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. The residue obtained after removal of the solvent was purified by column chromatography using dichloromethane/petroleum ether 1:2 for elution. The porphyrin fraction obtained was recrystallized from dichloromethane/methanol, to give 0.30 g (12% yield) of the title compound; MS m/e 1977.3 <sup>1</sup>H-NMR (d-TFA/CDCl<sub>3</sub>) δ ppm: -0.80 (br, NH), 1.5-3.1 (br, 80H, BH), 2.31 (s, 24H, CH<sub>3</sub>), 3.91 (s, 16H, CH<sub>2</sub>), 7.72 (s, 4H), 8.33 (s, 8H), 8.74 (s, 8H, β-H).

**[0078]      Compounds 29-31:**

The synthesis of compounds 29-31 is depicted in Scheme 5.

**[0079]      Synthesis of Compound 33:**

**5,15-bis[bis-3,5-(1-methyl-o-carboranyl)methylphenyl]porphyrin (33):** A solution of aldehyde **27** (0.446 g) and dipyrromethane **32** (0.146 g) in dry dichloromethane (100 mL) was purged with argon for 15 minutes and cooled to 0° C. TFA was added to the solution, and the final mixture was stirred at 0° C for 2 hours and then at room temperature overnight. After oxidation with *p*-chloranil (0.277 g, 1.13 mmol) for 6 hours at room temperature, the final reaction mixture was washed once with a saturated aqueous

NaHCO<sub>3</sub> solution, once with water, and once with brine before being dried over anhydrous MgSO<sub>4</sub>. The residue obtained after removal of the solvent was purified by column chromatography (alumina) using dichloromethane for elution. The porphyrin fraction obtained was recrystallized from acetone to give 33.6 % yield (0.192 g) of the title compound; MS m/e 1144.0; <sup>1</sup>H-NMR (d-TFA/CDCl<sub>3</sub>) δ ppm: -1.92 (br, NH), 1.4-3.2 (br, 40H, BH), 2.30 (s, 12H, CH<sub>3</sub>), 3.89 (s, 8H, CH<sub>2</sub>), 7.71 (s, 2H), 8.34 (s, 4H), 9.03 (d, 4H, β-H, J = 4.5 Hz), 9.61 (d, 4H, β-H, J = 4.5 Hz), 10.98 (s, 2H, *meso*-H).

**[0080]      Synthesis of Compound 34:**

**5,15-bis[bis-3,5-(1-methyl-*nido*-carboranyl)methylphenyl]porphyrin tetrapotassium salt (34):** Porphyrin **33** (0.100 g) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, the residue was re-dissolved in a 60% aqueous acetone solution and passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% acetone aqueous solution and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.102 g (92.8 % yield) of the title compound. <sup>1</sup>H-NMR (CD<sub>3</sub>COCD<sub>3</sub>) δ ppm: -2.84 (s, 2H, NH), -2.45 to -1.85 (br, 4H, BH), 0.9-2.4 (br, 32H, BH), 1.66 (s, 12H, CH<sub>3</sub>), 3.52 (s, 8H, CH<sub>2</sub>), 7.67 (s, 1H, ArH), 7.74 (s, 1H, ArH), 8.37 (s, 4H, ArH), 9.53 (dd, 4H, β-H), 9.61 (dd, 4H, β-H), 10.58 (s, 2H, *meso*-H). UV-Vis (acetone) λ<sub>max</sub>: 406 nm (ε 312,600), 502 (13,400), 536 (7,800), 576 (6,100), 630 (3,100).

**[0081]      Synthesis of Compound 35:**

**Zn(II)-5,15-bis[bis-3,5-(1-methyl-*o*-carboranyl)methylphenyl]porphyrin (35):** To a solution of porphyrin **33** (0.065 g, 0.057 mmol) in dichloromethane (100 mL) and THF (10 mL) was added ZnCl<sub>2</sub>•2H<sub>2</sub>O (0.031 g, 0.288 mmol), and the final mixture was stirred at

room temperature under argon overnight. The mixture was then washed once with water, dried over anhydrous  $\text{Na}_2\text{SO}_4$ , and the solvent was evaporated under vacuum. The residue was purified by column chromatography (dichloromethane/cyclohexane 2:1), and the pink-colored fraction was collected and recrystallized from dichloromethane/methanol to give 0.061 g (89 % yield) of the title compound.  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$  ppm: 1.4-3.0 (br, 40H, BH), 2.11 (s, 12H,  $\text{CH}_3$ ), 3.64 (s, 8H,  $\text{CH}_2$ ), 7.33 (s, 2H, ArH), 7.98 (s, 4H, ArH), 8.98 (d, 4H,  $\beta$ -H,  $J = 4.5$  Hz), 9.41 (d, 4H,  $\beta$ -H,  $J = 4.5$  Hz), 10.25 (s, 2H, *meso*-H).

**[0082]      Synthesis of Compound 36:**

**Zn(II)-5,15-bis[bis-3,5-(1-methyl-*nido*-carboranyl)methylphenyl]porphyrin tetrapotassium salt (36):** The Zn(II) complex **35** (0.050 g) was dissolved in a 3:1 mixture of pyridine and piperidine (4.0 mL), and was stirred at room temperature in the dark for 36 hours under argon. The solvent was completely removed under vacuum, the residue was re-dissolved in a 60% aqueous acetone solution, and was passed slowly through a Dowex 50W2-100 resin in potassium form. The porphyrin fraction was collected, dried under vacuum, re-dissolved in a 30% aqueous acetone solution and again passed through the ion-exchange resin. After removal of the solvent under vacuum, the tetraanionic porphyrin was recrystallized from methanol/diethyl ether, yielding 0.052 g (94.8 % yield) of the title compound;  $^1\text{H-NMR}$  ( $\text{CD}_3\text{COCD}_3$ )  $\delta$  ppm: -2.45 to -1.85 (br, 4H, BH), 0.9-2.4 (br, 32H, BH), 1.65 (s, 12H,  $\text{CH}_3$ ), 3.50 (s, 8H,  $\text{CH}_2$ ), 7.66 (s, 1H, ArH), 7.73 (s, 1H, ArH), 8.28 (s, 4H, ArH), 9.41 (dd, 4H,  $\beta$ -H), 9.47 (dd, 4H,  $\beta$ -H), 10.33 (s, 2H, *meso*-H). UV-Vis (acetone)  $\lambda_{\text{max}}$ : 412 nm ( $\epsilon$  263,100), 496 (1,400), 542 (9,900), 580 (1,070), MS (MALDI)  $m/e$  1319.4.

## ***In Vitro Analysis of Boronated Porphyrins***

### **[0083]        *Biological Studies***

**[0084]**        *In vitro* studies using rat 9L gliosarcoma cells, mouse B16 melanoma cells, hamster V79 fibroblast cells, and human U-373MG glioblastoma cells have been conducted. We found that all compounds studied had very low dark cytotoxicities, were readily taken up and retained by cells, and were localized in specific cell organelles, primarily those in close proximity to the cell nucleus.

**[0085]**        Our biological studies to date have indicated that these compounds have very low *in vivo* toxicities. So far we have determined maximum tolerated doses for compounds **4, 6, 16, 31**, and **33** using healthy female Balb/c mice. For all compounds tested, we found that the maximum tolerated dose exceeded 300 mg per kg of body weight.

### **[0086]        *Cytotoxicity / phototoxicity assays***

Human glioblastoma U373 MG, hamster V79 fibroblast, and mouse melanoma B16 cells were obtained from the American Type Culture Collection. Rat gliosarcoma cells were kindly provided by the University of California at San Francisco Brain Tumor Research Group. All cells were maintained in log phase monolayer cultures with RPMI 1640 supplemented with 10% fetal bovine serum and 2 mM glutamine.

**[0087]**        Cells were seeded in 96-well culture plates, and were allowed to settle and attach for 24-48 hours. Triplicate wells were then exposed to twofold serial dilutions of test compounds at concentrations up to 250  $\mu$ M. Compounds **4, 6, 10**, and **12** in crystalline form were carefully weighed and dissolved in 100% DMSO to prepare stock solutions. Subsequent dilutions were performed directly in the culture medium just before the medium was administered to cells. After short-term (2 hour) or long-term (24 hour) exposure, cells

were washed, and the wells were refilled with fresh culture medium. For dark toxicity trials, cells were allowed to proliferate for an additional 48-72 hours. For phototoxicity trials, washed cells were irradiated for 10 minutes with broad spectrum (600-700 nm) red light, and were then returned to an incubator for 48-72 hours.

**[0088]** Exposure to one of the compounds for 2 hours, followed by washing out the compound, did not inhibit proliferation of any of the three cell types. Exposure for 24 hours was inhibitory only at the higher concentrations ( $IC_{50} \geq 150 \mu M$ ) for 9L and U-373 MG cells, but B16 viability was unaffected. Whereas the metal-free porphyrins **4** and **10** displayed nearly identical  $IC_{50}$  values ( $\sim 150 \mu M$ ) in the affected cells, their Zn(II) complexes were about 20% less toxic ( $IC_{50} \sim 180-185 \mu M$ ).

**[0089]** We have also determined the cytotoxicity of porphyrins **4**, **16**, **18**, and **31** in V79 hamster fibroblast cells using a clonogenic assay. For each of these porphyrins, we found  $CS_{50} > 300 \mu M$ , and we observed no toxicity (i.e., no effect on colony survival) at concentrations up to  $200 \mu M$ .

**[0090]** In summary, these results showed that the *nido*-carboranyl compounds of the present invention display low dark toxicity.

#### **[0091] Binding Experiments**

The binding of several compounds to various ligands has been tested, and some of the resulting bound complexes have been examined in the solid state by X-ray diffraction. Figure 7 depicts the structure inferred from the X-ray data for compound **5** binding a methanol molecule axially. Figure 8 depicts the structure inferred from the X-ray data for compound **17** binding a water molecule axially. Figure 9 depicts the structure inferred from the X-ray data for compound **28** binding two chloride ions axially. Figure 10 depicts the structure inferred from the X-ray data for compound **28** binding oxygen atoms from two picrate ions axially. Figure 11 depicts the structure inferred from the X-ray data for compound **30** binding a pyridine molecule axially. Figure 12 depicts the structure

inferred from the X-ray data for compound **35** binding a pyridine molecule axially. The binding of pyridine, methanol, and biomolecules has also been examined in solution using UV-visible and fluorescence spectroscopy.

#### **[0092] Future Experiments**

Future planned experiments include the following:

- 1) Determine the binding affinity and binding kinetics for positively-charged molecules (e.g. pyridinium ions, bipyridinium ions, quaternary ammonium ions) to metal-free and metal-containing complexes of *nido*-carboranylporphyrins in solution. These studies will be followed by UV-Vis and fluorescence spectroscopy, mass spectrometry and nuclear magnetic resonance.
- 2) Determine the binding affinity and binding kinetics for neutral molecules (e.g. phenol, biphenols, amine-substituted molecules) to zinc and other metal-containing and metal-free complexes of *nido*- and *closo*-carboranylporphyrins in solution. These studies will be followed by UV-Vis and fluorescence spectroscopy, mass spectrometry and nuclear magnetic resonance.
- 3) Determine the binding affinity and binding kinetics of negatively-charged molecules (e.g. halides, phosphates, carboxylates) to metal-free or metal-containing, protonated or unprotonated, *nido*- and *closo*-carboranylporphyrins in solution. These studies will be followed by UV-Vis and fluorescence spectroscopy, mass spectrometry and nuclear magnetic resonance.

**[0093]** The complete disclosures of all references cited in this specification are hereby incorporated by reference. Also incorporated by reference are the complete disclosures of the following two provisional applications: Vicente and Marzilli, "Treating and



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preventing viral infections with porphyrin-based compounds," United States provisional patent application serial number 60/426,062, filed November 13, 2002; and Vicente, "Chelation of charged and uncharged molecules with porphyrin-based compounds," United States provisional patent application serial number 60/426,612, filed November 15, 2002. In the event of an otherwise irreconcilable conflict, however, the present specification shall control.